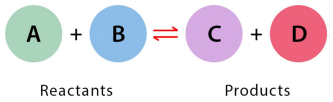
Dynamic Equilibrium

What is a reversible reaction?

Where products can react with each other

to produce the original reactants



Reaching Equilibrium

concentrations of reactants (A and B) fall

concentrations of products (C and D) will rise

eventually the forward reaction is going at the same rate as the reverse reaction

this is called dynamic equilibium

this can only happen in a closed system

i.e. no escape of reactants and products and nothing else can get in

Position of Equilibrium

Equilibrium does not mean the amounts of the products and reactants are equal

Equilibrium can be left or right !

Right : concentration of products greater than concentration of reactants

Left : Concentration of reactants greater than concentration of products

Changing the position of Equilibrium

Three factors can change the position of equilibrium

1. Temperature
2. Pressure
3. Concentration

What about catalysts?

They increase the rate that a reaction reaches equilibrium

BUT they do not change the position of the equilibrium